



Topic/Objective CHAPTER: 3

NAME:

Chemistry

Pd: 1 2 4 5 other

DATE

3/1/11

Essential Question

How do I Draw elements?

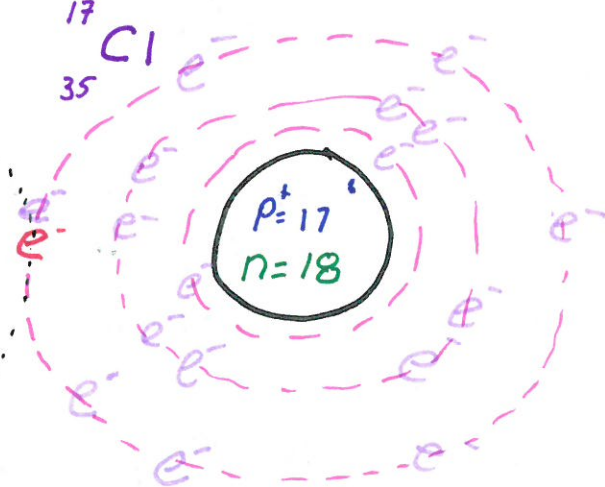
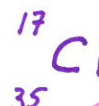
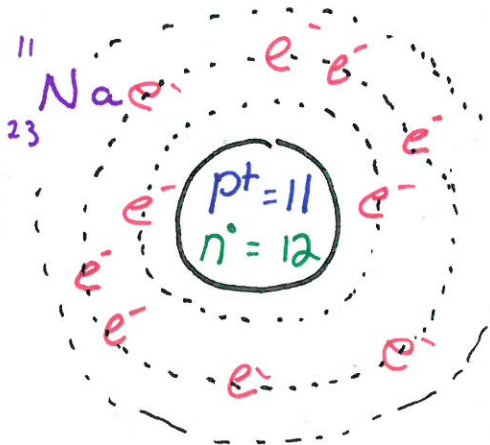
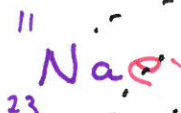
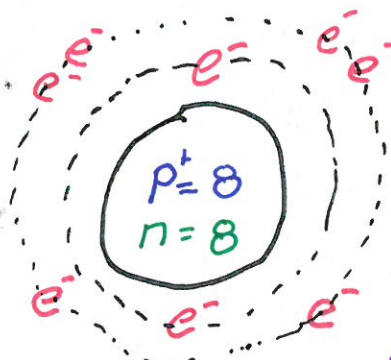
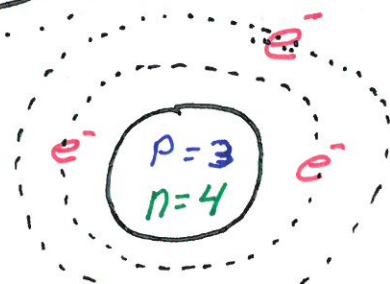
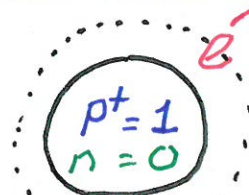
pt 3

Cue: Review:
Thoughts: Main Idea

NOTE Taking AREA:

Drawing elements

↳ Let's draw a few shells



Na wants to lose
1 e⁻ in the last
shell

Cl wants to GAIN
1 e⁻ in its last shell

NOTES CONTINUE ON OTHER SIDE

Positive 1⁺

Negative 1⁻

opposite charges



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Cue: Review:
Thoughts: Main Idea

NOTE Taking AREA:

most important level/shell

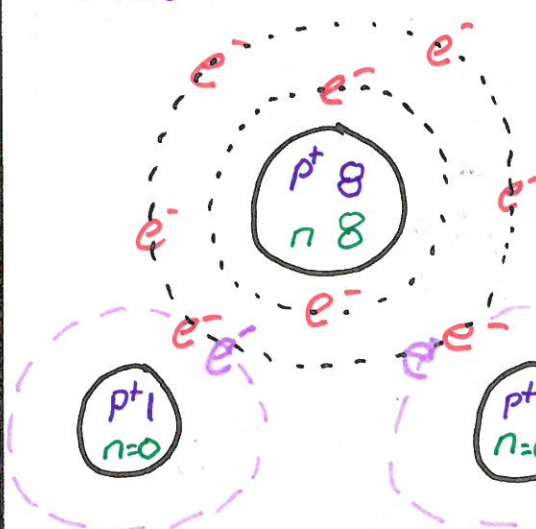
↳ The LAST orbital or Shell is the most important. It tells you WHAT the element wants to do. [IT's Behavior]

Valance Shell

↳ This Last Shell is called: Valance Shell

↳ is the # of electrons in the outer most shell

draw Water



↳ Valance shell of Oxygen is $6e^-$

↳ "L" shell would like to have $8e^-$ to be stable

↳ Hydrogen only has $1e^-$ in its Valance shell

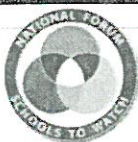
↳ the "K" shell want $2e^-$ to be stable

↳ So the Hydrogen Atom is Sharing its electron. w/o its e^- it would not be Hydrogen

SUMMARY:

↳ Oxygen is a 2^- charge
Hydrogen is a 1^+ charge
So you need 2 Hydrogen elements to make Oxygen stable





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Chemistry

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Pt 2

Essential Question

What are the Elements of the Periodic Table?

Cue: Review:
Thoughts: Main Idea

NOTE Taking AREA:

Symbol

↳ Capital letter
↳ sometimes a lower case letter has to follow it.

State of matter

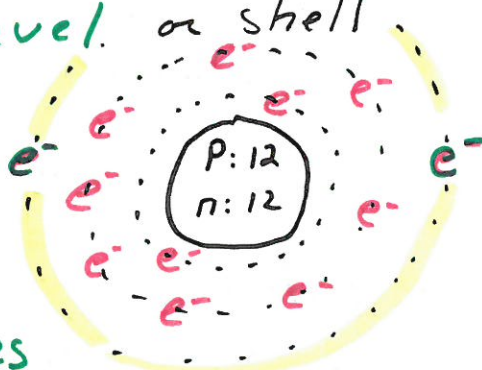
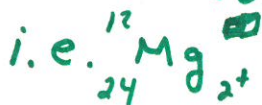
↳ How an element exist @ room temperature

from Book
page = 73

↳ H₂O [water] is unique to ⊕ b/c it exist in ALL 3 states @ the same time.

of Valence electrons

↳ # of e⁻ in the outer most energy level or shell



↳ This determines if the element will:

Transfer e⁻ or Share e⁻

↳ Ionic

↳ opp. charges

↳ Covalent

NOTES CONTINUE ON OTHER SIDE